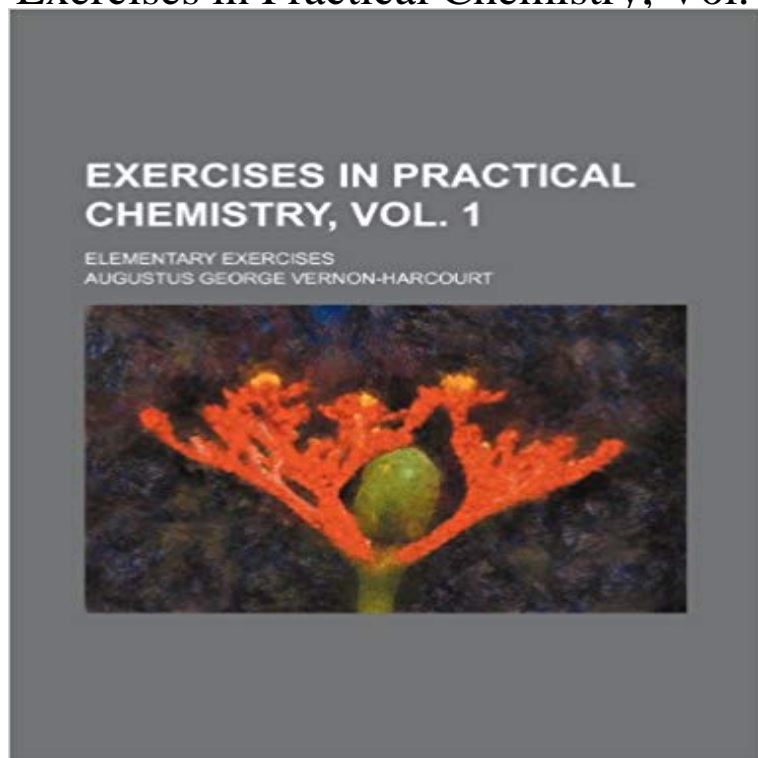


Exercises in Practical Chemistry, Vol. 1 Volume 1; Elementary Exercises



This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1880 Excerpt: ... SECTION II. PREPARATION AND EXAMINATION OF THE PROPERTIES OF THE PRINCIPAL NON-METALLIC RADICLES AND THEIR COMPOUNDS. An asterisk is prefixed to those experiments which illustrate properties of the substance which are considered especially applicable for its detection in the course of qualitative analysis. 1. OXYGEN and OXIDES. Apparatus required--Pneumatic trough; argand burner; fish-tail burner, or spirit lamp; Bunsen burner; retort stand; sand-bath; piece of wire gauze; corks; cork-borers; rat-tail file; three-square file; Florence flask; glass tubing, about 5 mm. in external diameter; porcelain mortar; two small porcelain dishes; scales and weights; sheets of writing-paper; deflagrating jar; deflagrating cup; one wide-mouthed stoppered bottle, holding about 700 c.c.; two ditto, holding about 20 c.c.; two cylindrical gas jars, 20 x 5 cm.; ground-glass disc, 8 cm. in diameter; taper on wire (Fig 9); cedar matches; piece of watch-spring; German tinder, or fusee; crucible tongs; jug of water; blotting-paper; potassium chlorate; manganese dioxide; clean sand; sulphur; box of test-papers. Oxygen is generally obtained by the action of heat on some compound containing it. Thus when mercury oxide was heated (p. 69) it was found to give off oxygen; and this method is of interest as being the one used by Priestley in his discovery of oxygen. But practically another substance, potassium chlorate, is nearly always used in the preparation of the gas. This salt contains potassium chlorine and oxygen, and when it is heated the whole of the oxygen in it (about one-third of its weight) is evolved,

and a compound of potassium and chlorine (potassium chloride) remains. This may be tried on a small scale as follows. $2\text{KClO}_2 = 2\text{KCl} + 3\text{O}_2$. Place a few crystals of potassiu...

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